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5/15 - Energetics - IB Chemistry - Google Sites

IB Chemistry SL: Energetics [Topic 5] by Ben Skirrow on Prezi. Bond enthalpy aka dissociation enthalpy -- the enthalpy change when one mole of bonds broken homolitically in the gas phase.

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15.1 Born Haber cycles Understandings: Enthalpy of solution, hydration enthalpy and lattice enthalpy are related in an energy cycle. Applications and skills: Construction of Born-Haber cycles for group 1 and 2 oxides and chlorides. Calculation of enthalpy changes from Born-Haber or dissolution energy cycles.

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15.2 Entropy and spontaneity presentation. 15.2_entropy_and_spontaneity.pdf:

Topic 5 and 15 - Energetics - Department of Natural

IB Chemistry standard level revision notes on energetics. 5.4 Bond enthalpy and bond dissociation enthalpy. Bond dissociation enthalpy is the enthalpy change when one mole of specific bonds are broken.. $X-Y(g) \rightarrow X(g) + Y(g)$: ΔH (dissociation).. Bond enthalpy is the average value of a particular type of bond which has been measured over a range of molecules.

IB Chemistry standard level revision notes: Energetics

15.2 - Born Haber Cycles Born Haber cycles are the application of Hess' law to ionic systems. An ionic solid consists of a giant structure of ions held together in a giant lattice.

IB Chemistry higher level revision notes: Energetics

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The IB Chemistry course at Scots College will provide opportunities for ... stressed. Beyond this, students will be better prepared for the IB exam with the ultimate goal of achieving the diploma. Chemistry, as all sciences, is an international endeavour. ... Topic 15: Energetics 15.1 Standard enthalpy changes of reaction 15.2 Born-Haber ...

IB Chemistry - Scots College School, Wellington

Overall IB Chemistry Reviews. IB Chemistry HL 31 Common Mistakes: Richard Thornley, the author of this video, has several other helpful videos on IB Chemistry SL and HL available for free on YouTube; IB Chemistry Web: This site goes over the syllabus in-depth and explains key definitions and facts you need to know.

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15. Energetics/thermochemistry 16. Chemical kinetics 17. Equilibrium 18. Acids and bases 19. Redox

processes 20. Organic chemistry 21. Measurement and analysis 60 2 4 7 7 6 4 10 6 12 2 Option (Choice of one out of four) A. Materials B. Biochemistry C. Energy D. Medicinal chemistry 25 25 25 25 25 Practical scheme of work

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15.1.1 Standard state is what the element naturally tends to be at. The standard enthalpy of formation is the change in enthalpy that accompanies the formation of one mole of a compound from its elements with all substances in their standard states.

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Chemistry i CHEMISTRY SYLLABUS, ASSESSMENT and UNIT PLANNERS ... 2014 Edition IB Chemistry study guide ,â,- ISBN: 978-0-19-839353-5 ... AHL: Energetics/thermochemistry (Topic 15) Syllabus aims addressed Core: The enthalpy changes from chemical reactions can be calculated from their effect on the temperature of their surroundings.

Chemistry i - utu

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